

# Answers For Chemistry Module 13

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The chemical equation contains no phases, so we must use bond energies to solve this problem. This means we start with Lewis structures. Using those and Table 13.1, we know that 2 Cl-Cl bonds

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(each worth 240 kJ/mole) and 4 H-O bonds (each worth 459 kJ/mole) must be broken while 4 H-Cl bonds (each worth 428 kJ/mole) and 1 O=O bond (worth 494 kJ/mole) must be formed.

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Apologia Chemistry Module 13 Vocabulary. Change in Enthalpy. Hess's Law. Enthalpy of Formation. Standard Enthalpy of Formation. The energy change that accompanies a chemical reaction. Enthalpy is a state function. The change of H of a formation reaction. Change of H of a formation reaction at T= 298K and P= 1.00 atm.

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